

# Ionic Bonding with Lewis Structures

Step:

1. Draw Lewis Structure for each element.
  2. Decide if elements will gain or lose electrons.
  3. One element will full fill octet rule (8 electrons), other will lose all.
  4. Duplicate metal or nonmetal as needed so each nonmetal will get to "8" electrons.
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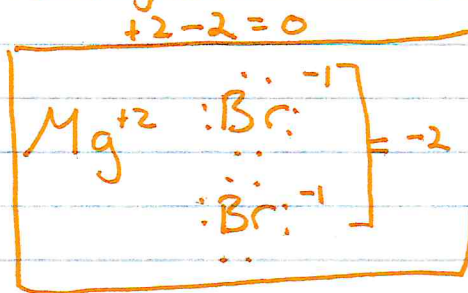
5. Write number of metal atoms in box (no valence electrons)  
lose electron  $\rightarrow$  get positive
6. Write nonmetal Lewis Structure/electrons with charge.
7. Charges of non metal and metal should equal 0, zero.

Problem #3

magnesium bromide (magnesium + bromine)



Steps 1-4



MgBr<sub>2</sub>

Steps 5-7