

Complete the following table. You will need a Periodic Table to complete some portions of the table.

Atomic Number	Mass Number	Protons	Neutrons	Electrons	Nuclide Notation
7	14	7	7	7	$^{14}_7\text{N}$
16	32	16	16	16	$^{32}_{16}\text{S}$
22	48	22	26	22	$^{48}_{22}\text{Ti}$
36	84	36	48	36	$^{84}_{36}\text{Kr}$
86	220	86	134	86	$^{220}_{86}\text{Rn}$
53	113	53	60	53	$^{113}_{53}\text{I}$
31	70	31	39	31	$^{70}_{31}\text{Ga}$
20	42	20	22	20	$^{42}_{20}\text{Ca}$
34	78	34	44	34	$^{78}_{34}\text{Se}$
25	55	25	30	25	$^{55}_{25}\text{Mn}$
18	34	18	16	18	$^{34}_{18}\text{Ar}$

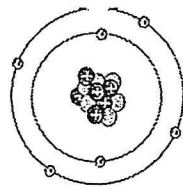
From periodic table
Mass - Atomic #

Practice Problems: Complete the following chart. You may need to reference a Periodic Table to complete some portions of the chart.

Element	Atomic Number	Protons	Electrons	Neutrons
I Do	9	9	9	10
We Do	12	12	12	12
You Do	56	56	56	81

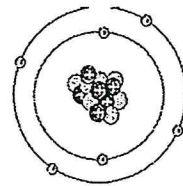
Remember, the proton name identify the type of atom.

Carbon-12



6 protons
6 neutrons
6 electrons

Carbon-13

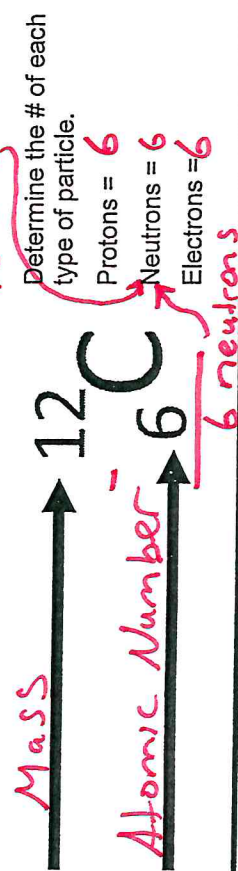


6 protons
7 neutrons
6 electrons

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Isotopes is an element where the protons are the same but neutrons are different. So mass is different too.

The composition of any atom located on the Periodic Table can be represented in a shorthand form called isotope notation.



Practice: Determine the number of each type of subatomic particle.

