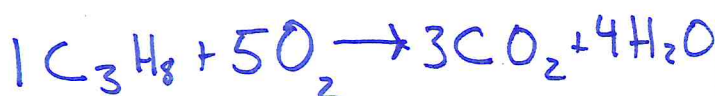
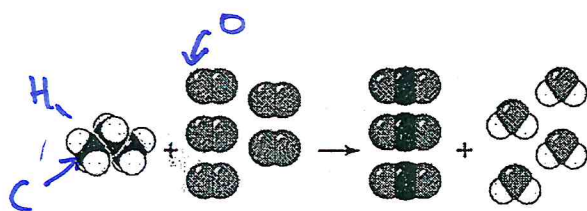
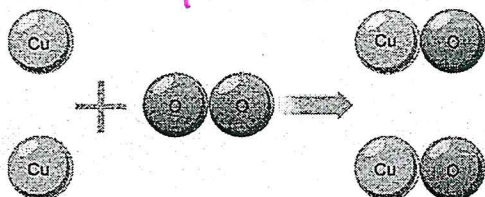


Date:

What is the mole ratio for each equation below?

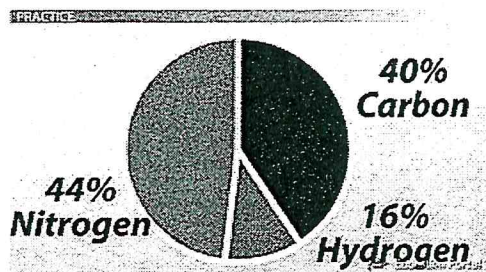
Synthesis



Combustion

Date:

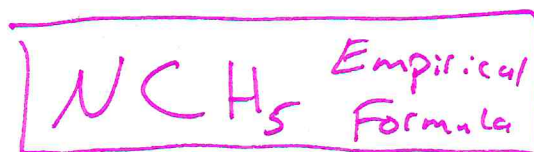
Calculate the empirical formula of the compound below.



$$44\text{g N} \times \frac{1\text{mol N}}{14\text{g N}} = \frac{3.14}{3.14} = 1$$

$$40\text{g C} \times \frac{1\text{mol C}}{12\text{g C}} = \frac{3.33}{3.14} = 1$$

$$16\text{g H} \times \frac{1\text{mol H}}{1.01\text{g H}} = \frac{15.84}{3.14} = 5$$



Date:

Write the balanced equation for the reaction below.

~~H^+~~ ~~F^-~~ + ~~Ca^{2+}~~ ~~OH^-~~
Hydrogen Fluoride + Calcium Hydroxide $\rightarrow$



Acid

Base



water

What type of reaction is this?

Double Replacement

$\rightarrow$ Neutralization

Date:

List the phase changes that are exothermic.

Exit (heat, energy)
Release

Freezing: liquid to Solid

Condensation: gas to liquid

Deposition: gas to Solid