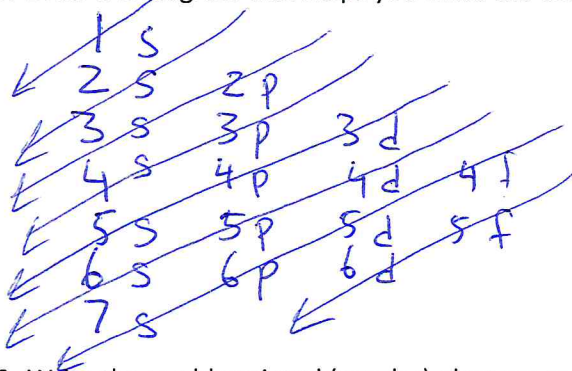


Name: Key Date: _____ Class: _____ Hour: _____

15. Draw the diagram that helps you write the correct electron configuration.



16. Write the unabbreviated (regular) electron configurations of the following elements:

- a. Nitrogen $1s^2 2s^2 2p^3$
- b. Phosphorus $1s^2 2s^2 2p^6 3s^2 3p^3$
- c. Potassium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$
- d. Titanium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^2$

17. Write the abbreviated (noble gas) electron configurations of the following elements:

- a. Beryllium $[Ar] 4s^2 3d^{10} 4p^5$
- b. Zinc $[Ar] 4s^2 3d^{10}$
- c. Gallium $[Ar] 4s^2 3d^{10} 4p^1$
- d. Sulfur $[Ne] 3s^2 3p^4$

18. The following electron configurations belong to which elements:

- a. $1s^2 2s^2 2p^6 3s^1$ Na Sodium
- b. $[Xe] 6s^2 4f^{14} 5d^{10} 6p^2$ Pb Lead
- c. $[Kr] 5s^2 4d^{10}$ Cd Cadmium
- d. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10}$ Zn Zinc